Acta Crystallographica Section E
Structure Reports Online
ISSN 1600-5368

2-Chloro-*N*-isopropyl-*N*-phenylacetamide

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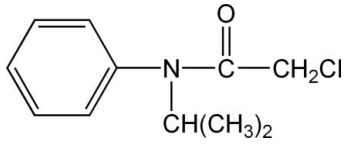
Received 26 June 2010; accepted 8 July 2010

Key indicators: single-crystal X-ray study; T = 298 K; mean σ(C–C) = 0.003 Å; R factor = 0.038; wR factor = 0.111; data-to-parameter ratio = 15.4.

In the title compound, $C_{11}H_{14}ClNO$, the herbicide propachlor, there are significant differences between the three N—C bond lengths [N—$C_{carbonyl}$ = 1.354 (3) Å, N—C_{phenyl} = 1.444 (2) Å and N—$C_{isopropyl}$ = 1.496 (3) Å], indicating the presence of π delocalization involving the carbonyl group. The N atom lies 0.074 (2) Å from the plane defined by the the three bonded C atoms.

Related literature

For studies of propachlor and its derivatives, see: Dhillon & Anderson (1972); Kleudgen (1980).



Experimental

Crystal data

$C_{11}H_{14}ClNO$
M_r = 211.68
Monoclinic, $P2_1/n$
a = 11.9190 (11) Å
b = 7.8042 (8) Å
c = 12.3789 (13) Å
β = 98.963 (1)°
V = 1137.4 (2) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.30 mm^{-1}
T = 298 K
0.47 × 0.45 × 0.44 mm

Data collection

Bruker APEXII CCD diffractometer
Absorption correction: multi-scan (*SADABS*; Bruker, 2005)
T_{min} = 0.870, T_{max} = 0.878
5426 measured reflections
2004 independent reflections
1501 reflections with $I > 2\sigma(I)$
R_{int} = 0.025

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.038
$wR(F^2)$ = 0.111
S = 1.03
2004 reflections
130 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.21 e Å^{-3}
$\Delta\rho_{min}$ = −0.19 e Å^{-3}

Data collection: *SMART* (Bruker, 1997); cell refinement: *SAINT* (Bruker, 1997); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

This work was supported by the PhD Programs Foundation of the Ministry of Education of China (No. 20090204120033)

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: ZS2048).

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Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.

supplementary materials

Acta Cryst. (2010). E**66**, o2017 [**doi:10.1107/S1600536810027157**]

2-Chloro-*N*-isopropyl-*N*-phenylacetamide

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Comment

Propachlor (2-chloro-*N*-isopropyl-*N*-phenylacetamide) and its derivatives have been widely studied as a pre-emergent herbicide used to control broadleaf weeds and grasses (Dhillon *et al.*, 1972; Kleudgen, 1980). Propachlor may also be used as a precursor in the synthesis of indole-2-one compounds and in the course of exploring new indole-2-one compounds, we synthesized the title compound $C_{11}H_{14}ClNO$ (I), the structure of which is reported here.

In structure of (I) (Fig. 1), there are obvious differences between the three C—N bond lengths (N—$C_{carbonyl}$, 1.354 (3) Å; N—C_{phenyl}, 1.444 (2) Å; N—$C_{isopropyl}$, 1.496 (3) Å, indicating the presence of π delocalization involving the carbonyl group. Also N1 lies close to the plane defined by the three bonded carbon atoms C1, C3 and C9 [0.074 (2) Å].

As expected, there are no classic hydrogen bonds in the structure (Fig. 2). However, there is a weak intermolecular aliphatic C11—H11A···$O1^{i}$ interaction [symmetry code: (i) -*x* + 1/2, *y* - 1/2, -*z* + 1/2] stabilizing the packing. This intermolecular hydrogen bond is characterized by the parameters 0.96 Å (C11—H11A) and 2.56 Å (H11A···$O1^{i}$).

Experimental

N-Isopropylbenzenamine (1.00 g, 7.41 mmol) was dissolved in toluene (5.0 mL) and cooled to 273 K, after which a solution of 2-chloroacetyl chloride (0.90 g, 8.03 mmol) in toluene (5.0 mL) was slowly added over 0.5 h. with stirring. The mixture was then refluxed for 2 h. then slowly cooled to room temperature over 8 h. Colorless block crystals of (I) were formed (1.33 g, yield 85%).

Refinement

All H atoms were placed in geometrically calculated positions and refined using a riding model with C—$H_{aromatic}$ = 0.93 %A and C—$H_{aliphatic}$ = 0.96–0.97 %A, with $U_{iso} = 1.2U_{eq}$(C), or $1.5U_{eq}$(C) for CH_3 groups.

Figures

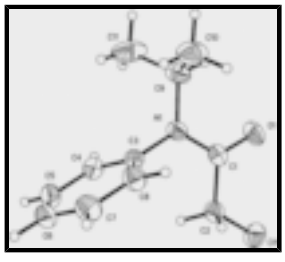

Fig. 1. The molecular comformation and atom numbering scheme for (I). Displacement ellipsoids are drawn at the 30% probability level.

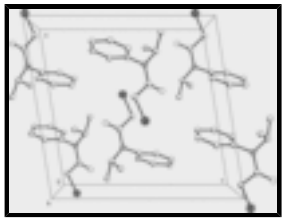

Fig. 2. The molecular packing of (I) viewed down the *b* axis of the unit cell.

2-Chloro-*N*-isopropyl-*N*-phenylacetamide

Crystal data

$C_{11}H_{14}ClNO$	$F(000) = 448$
$M_r = 211.68$	$D_x = 1.236$ Mg m^{-3}
Monoclinic, $P2_1/n$	Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Hall symbol: -P 2yn	Cell parameters from 2899 reflections
$a = 11.9190$ (11) Å	$\theta = 2.2$–$25.0°$
$b = 7.8042$ (8) Å	$\mu = 0.30$ mm^{-1}
$c = 12.3789$ (13) Å	$T = 298$ K
$\beta = 98.963$ (1)°	Block, colorless
$V = 1137.4$ (2) Å^3	0.47 × 0.45 × 0.44 mm
$Z = 4$	

Data collection

Bruker APEXII CCD diffractometer	2004 independent reflections
Radiation source: fine-focus sealed tube	1501 reflections with $I > 2\sigma(I)$
graphite	$R_{int} = 0.025$
φ and ω scans	$\theta_{max} = 25.0°$, $\theta_{min} = 2.2°$
Absorption correction: multi-scan (*SADABS*; Bruker, 2005)	$h = -14\rightarrow12$
$T_{min} = 0.870$, $T_{max} = 0.878$	$k = -9\rightarrow8$
5426 measured reflections	$l = -14\rightarrow14$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.038$	H-atom parameters constrained
$wR(F^2) = 0.111$	$w = 1/[\sigma^2(F_o^2) + (0.0469P)^2 + 0.4408P]$ where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.03$	$(\Delta/\sigma)_{max} < 0.001$
2004 reflections	$\Delta\rho_{max} = 0.21$ e Å^{-3}
130 parameters	$\Delta\rho_{min} = -0.19$ e Å^{-3}

0 restraints

Primary atom site location: structure-invariant direct methods

Extinction correction: *SHELXL97* (Sheldrick, 2008), $Fc^*=kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$

Extinction coefficient: 0.079 (5)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
Cl1	0.57852 (5)	0.22476 (9)	0.44873 (5)	0.0711 (3)
N1	0.25393 (14)	0.3526 (2)	0.40187 (13)	0.0483 (4)
O1	0.40481 (14)	0.4455 (2)	0.32694 (12)	0.0662 (5)
C1	0.36637 (17)	0.3629 (3)	0.39670 (16)	0.0475 (5)
C2	0.44299 (18)	0.2648 (3)	0.48500 (18)	0.0543 (6)
H2A	0.4074	0.1566	0.4980	0.065*
H2B	0.4520	0.3301	0.5525	0.065*
C3	0.21346 (16)	0.2664 (3)	0.49159 (15)	0.0438 (5)
C4	0.17401 (18)	0.0998 (3)	0.47994 (18)	0.0548 (6)
H4	0.1734	0.0423	0.4140	0.066*
C5	0.1354 (2)	0.0186 (3)	0.5666 (2)	0.0674 (7)
H5	0.1097	−0.0939	0.5592	0.081*
C6	0.1349 (2)	0.1042 (4)	0.6635 (2)	0.0691 (7)
H6	0.1073	0.0504	0.7211	0.083*
C7	0.1751 (2)	0.2691 (4)	0.67528 (19)	0.0701 (7)
H7	0.1757	0.3260	0.7414	0.084*
C8	0.21471 (19)	0.3513 (3)	0.58983 (17)	0.0586 (6)
H8	0.2420	0.4629	0.5983	0.070*
C9	0.1714 (2)	0.4559 (3)	0.32459 (18)	0.0641 (7)
H9	0.2083	0.4850	0.2615	0.077*
C10	0.1415 (3)	0.6210 (3)	0.3759 (3)	0.0939 (10)
H10A	0.1033	0.5964	0.4368	0.141*
H10B	0.0926	0.6875	0.3228	0.141*
H10C	0.2097	0.6845	0.4006	0.141*
C11	0.0654 (3)	0.3538 (4)	0.2827 (3)	0.1013 (11)
H11A	0.0865	0.2471	0.2530	0.152*
H11B	0.0186	0.4179	0.2267	0.152*
H11C	0.0240	0.3318	0.3418	0.152*

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cl1	0.0516 (4)	0.0819 (5)	0.0834 (5)	0.0023 (3)	0.0212 (3)	−0.0015 (3)
N1	0.0526 (10)	0.0511 (10)	0.0426 (9)	0.0058 (8)	0.0115 (7)	0.0076 (8)
O1	0.0760 (11)	0.0647 (11)	0.0635 (9)	−0.0032 (8)	0.0288 (8)	0.0168 (8)
C1	0.0592 (13)	0.0403 (11)	0.0455 (11)	−0.0013 (9)	0.0157 (10)	−0.0021 (9)
C2	0.0477 (12)	0.0591 (14)	0.0588 (13)	0.0012 (10)	0.0167 (10)	0.0089 (11)
C3	0.0410 (10)	0.0493 (12)	0.0420 (10)	0.0039 (9)	0.0091 (8)	0.0049 (9)
C4	0.0519 (12)	0.0549 (14)	0.0582 (13)	−0.0021 (10)	0.0106 (10)	−0.0024 (11)
C5	0.0572 (14)	0.0606 (15)	0.0859 (17)	−0.0076 (11)	0.0157 (12)	0.0164 (14)
C6	0.0581 (14)	0.090 (2)	0.0622 (15)	−0.0004 (13)	0.0178 (11)	0.0268 (15)
C7	0.0761 (17)	0.091 (2)	0.0456 (13)	−0.0016 (14)	0.0178 (12)	0.0005 (13)
C8	0.0685 (14)	0.0576 (14)	0.0514 (12)	−0.0041 (11)	0.0150 (10)	−0.0015 (11)
C9	0.0692 (15)	0.0715 (17)	0.0521 (12)	0.0177 (12)	0.0110 (11)	0.0201 (12)
C10	0.095 (2)	0.0512 (16)	0.122 (2)	0.0114 (15)	−0.0266 (18)	0.0041 (16)
C11	0.112 (2)	0.078 (2)	0.094 (2)	0.0226 (18)	−0.0480 (18)	−0.0207 (17)

Geometric parameters (Å, °)

Cl1—C2	1.771 (2)	C6—C7	1.374 (4)
N1—C1	1.354 (3)	C6—H6	0.9300
N1—C3	1.444 (2)	C7—C8	1.381 (3)
N1—C9	1.496 (3)	C7—H7	0.9300
O1—C1	1.222 (2)	C8—H8	0.9300
C1—C2	1.518 (3)	C9—C10	1.503 (4)
C2—H2A	0.9700	C9—C11	1.515 (4)
C2—H2B	0.9700	C9—H9	0.9800
C3—C4	1.383 (3)	C10—H10A	0.9600
C3—C8	1.383 (3)	C10—H10B	0.9600
C4—C5	1.385 (3)	C10—H10C	0.9600
C4—H4	0.9300	C11—H11A	0.9600
C5—C6	1.373 (4)	C11—H11B	0.9600
C5—H5	0.9300	C11—H11C	0.9600
C1—N1—C3	121.03 (16)	C6—C7—C8	120.6 (2)
C1—N1—C9	119.66 (17)	C6—C7—H7	119.7
C3—N1—C9	118.53 (16)	C8—C7—H7	119.7
O1—C1—N1	123.2 (2)	C7—C8—C3	119.5 (2)
O1—C1—C2	121.67 (19)	C7—C8—H8	120.2
N1—C1—C2	115.15 (16)	C3—C8—H8	120.2
C1—C2—Cl1	112.14 (14)	N1—C9—C10	111.53 (19)
C1—C2—H2A	109.2	N1—C9—C11	111.5 (2)
Cl1—C2—H2A	109.2	C10—C9—C11	110.8 (2)
C1—C2—H2B	109.2	N1—C9—H9	107.6
Cl1—C2—H2B	109.2	C10—C9—H9	107.6
H2A—C2—H2B	107.9	C11—C9—H9	107.6
C4—C3—C8	119.95 (19)	C9—C10—H10A	109.5

C4—C3—N1	120.46 (18)	C9—C10—H10B	109.5
C8—C3—N1	119.59 (19)	H10A—C10—H10B	109.5
C3—C4—C5	119.9 (2)	C9—C10—H10C	109.5
C3—C4—H4	120.1	H10A—C10—H10C	109.5
C5—C4—H4	120.1	H10B—C10—H10C	109.5
C6—C5—C4	120.1 (2)	C9—C11—H11A	109.5
C6—C5—H5	119.9	C9—C11—H11B	109.5
C4—C5—H5	119.9	H11A—C11—H11B	109.5
C5—C6—C7	120.0 (2)	C9—C11—H11C	109.5
C5—C6—H6	120.0	H11A—C11—H11C	109.5
C7—C6—H6	120.0	H11B—C11—H11C	109.5
C3—N1—C1—O1	−174.27 (19)	N1—C3—C4—C5	−179.95 (19)
C9—N1—C1—O1	−4.6 (3)	C3—C4—C5—C6	0.8 (3)
C3—N1—C1—C2	5.3 (3)	C4—C5—C6—C7	−1.5 (4)
C9—N1—C1—C2	174.98 (19)	C5—C6—C7—C8	1.0 (4)
O1—C1—C2—Cl1	−20.0 (3)	C6—C7—C8—C3	0.1 (4)
N1—C1—C2—Cl1	160.48 (15)	C4—C3—C8—C7	−0.8 (3)
C1—N1—C3—C4	−98.8 (2)	N1—C3—C8—C7	179.5 (2)
C9—N1—C3—C4	91.4 (2)	C1—N1—C9—C10	−96.3 (2)
C1—N1—C3—C8	81.0 (3)	C3—N1—C9—C10	73.7 (3)
C9—N1—C3—C8	−88.8 (2)	C1—N1—C9—C11	139.2 (2)
C8—C3—C4—C5	0.3 (3)	C3—N1—C9—C11	−50.8 (3)

Fig. 1

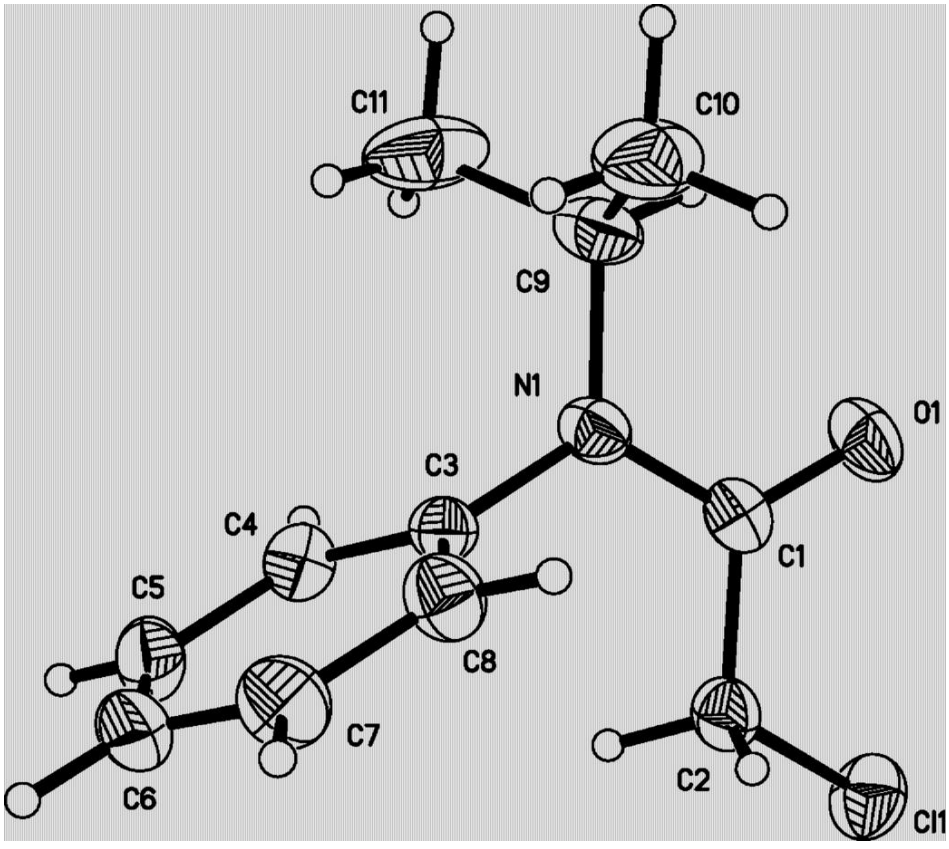

Fig. 2